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in the reduction products, and the degree of reducibility $A =$ percentage of Ni reduced, from 100 to 1,000°. From the fact that Ni is reduced at as low as 300° at which temperature only NiO, not Ni silicates, are reducible, indicates that even in silicate ores Ni is present not as silicate; this is borne out by the often encountered ease of leaching Ni out from these ores with acids. The temperatures of complete extraction of Ni by reduction with H_2 and with CO from both silicate and oxidized ores are given. The effect of duration of the reduction process depends on the sort of ore and on temperature. Ni obtained by reduction with H_2 possesses a particularly high reactivity in hydrometallurgical applications, in contrast to Ni reduced by CO and C; low reduction temperature is favorable from this viewpoint.

"Solubility in the System $CaCO_3$ - $CaSO_4$ - $NaCl$ - CO_2 - H_2O at 25°," E. B. Shternina, E. V. Frolova, Inst Gen and Inorg Chem, Acad Sci USSR

"Compt Rend Acad Sci URSS" Vol 47, 1945, pp 33-5
"Dok Akad Nauk SSSR" Vol 47, 1945, pp 34-6,

The solubilities of $CaCO_3$ and $CaSO_4$ in water containing 0-2% NaCl at CO_2 pressures of 0.0013-1.0 atmosphere are given. The data are expressed graphically in terms of the $Ca(HCO_3)_2$ and $CaSO_4$ contents of the solutions at equilibrium. Use of the results to predict solution of calcite and gypsum in river waters is indicated.

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